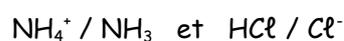
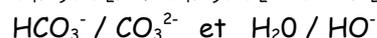
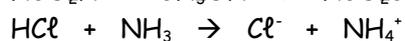
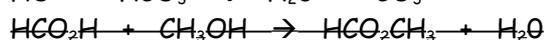
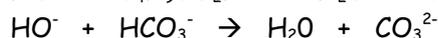
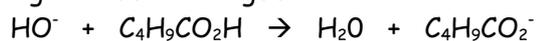
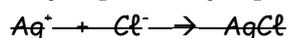
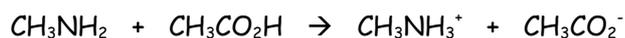


## Chapitre 7. Réactions acido-basiques Exemples de cours - corrigé

### Exemple n°1 : Réaction entre 2 couples acide/base

- l'acide nitreux  $\text{HNO}_2$                        $\text{HNO}_2 = \text{NO}_2^- + \text{H}^+$   
l'ammoniac  $\text{NH}_3$ .                               $\text{NH}_4^+ = \text{NH}_3 + \text{H}^+$
- $\text{HNO}_2(\text{aq}) + \text{NH}_3(\text{aq}) \rightarrow \text{NO}_2^-(\text{aq}) + \text{NH}_4^+(\text{aq})$

### Exemple n° 2 : Reconnaître une réaction acido-basique



### Exemple n°3 : L'ion hydrogénocarbonate

- $\text{HCO}_3^- = \text{CO}_3^{2-} + \text{H}^+$
- $\text{CO}_2, \text{H}_2\text{O} = \text{HCO}_3^- + \text{H}^+$
- Car il existe à la fois en tant qu'acide et que base
- $\text{CO}_2(\text{aq}), \text{H}_2\text{O} + \text{CO}_3^{2-}(\text{aq}) \rightarrow 2 \text{HCO}_3^-(\text{aq})$

### Exemple n°4 : pH

$[\text{H}_3\text{O}^+]$	$6,0 \cdot 10^{-5}$	$3,9 \cdot 10^{-4}$	$5,4 \cdot 10^{-8}$	$1,26 \cdot 10^{-4}$	$1,58 \cdot 10^{-7}$	$6,31 \cdot 10^{-12}$
pH	4,2	3,4	7,3	3,9	6,8	11,2